

AEDC-TDR-63-139

6342

409 870

CATALOGED BY DDC
AS AD 1000



**CHARTS OF SPHERE STAGNATION HEAT-TRANSFER RATE
IN AIR AND NITROGEN AT HIGH TEMPERATURES**

By

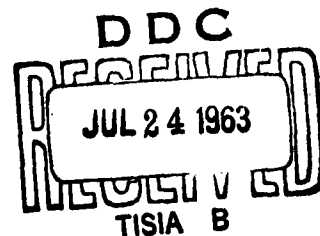
Clark H. Lewis
von Kármán Gas Dynamics Facility
and

Ernest G. Burgess, III
Scientific Computing Services
Office of the Director of Engineering
ARO, Inc.

TECHNICAL DOCUMENTARY REPORT NO. AEDC-TDR-63-139

July 1963

AFSC Program Area 040A



(Prepared under Contract No. AF 40(600)-1000 by ARO, Inc.,
contract operator of AEDC, Arnold Air Force Station, Tenn.)

**ARNOLD ENGINEERING DEVELOPMENT CENTER
AIR FORCE SYSTEMS COMMAND
UNITED STATES AIR FORCE**

NOTICES

Qualified requesters may obtain copies of this report from ASTIA. Orders will be expedited if placed through the librarian or other staff member designated to request and receive documents from ASTIA.

When Government drawings, specifications or other data are used for any purpose other than in connection with a definitely related Government procurement operation, the United States Government thereby incurs no responsibility nor any obligation whatsoever; and the fact that the Government may have formulated, furnished, or in any way supplied the said drawings, specifications, or other data, is not to be regarded by implication or otherwise as in any manner licensing the holder or any other person or corporation, or conveying any rights or permission to manufacture, use, or sell any patented invention that may in any way be related thereto.

CHARTS OF SPHERE STAGNATION HEAT-TRANSFER RATIO
IN AIR AND NITROGEN AT HIGH TEMPERATURES

By
Clark H. Lewis
von Kármán Gas Dynamics Facility
and
Ernest G. Burgess, III
Scientific Computing Services
Office of the Director of Engineering
ARO, Inc.
a subsidiary of Sverdrup and Parcel, Inc.

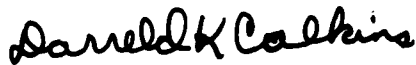
July 1963
ARO Project No. VT8002

ABSTRACT

Fay and Riddell's theory for laminar heat transfer at the stagnation point is applied to the sphere in equilibrium air and nitrogen. Charts are presented of the heat-transfer rate as a function of stagnation temperature and pressure or free-stream velocity and stagnation pressure in the pressure range between 0.1 and 10 atm and in the temperature range from 2000 to 15,000°K. The effects of Lewis number are considered. Comparisons between Fay and Riddell's theory, the recent theory of Fay and Kemp, and shock-tube experiments are presented.

PUBLICATION REVIEW

This report has been reviewed and publication is approved.



Darreld K. Calkins
Major, USAF
AF Representative, VKF
DCS/Test



Jean A. Jack
Colonel, USAF
DCS/Test

CONTENTS

	<u>Page</u>
ABSTRACT.	ii
NOMENCLATURE.	iv
1.0 INTRODUCTION	1
2.0 THERMODYNAMIC AND TRANSPORT PROPERTIES	
2.1 Air	2
2.2 Nitrogen.	2
3.0 RESULTS AND DISCUSSION	3
4.0 CONCLUDING REMARKS	5
REFERENCES	5

ILLUSTRATIONS

Figure

1. Viscosity Ratio for Air	8
2. Lewis Number Correlation for Air	9
3. Viscosity Ratio for Nitrogen	10
4. Stagnation Point Heat Transfer in Air ($Le = 1$)	11
5. Stagnation Point Heat Transfer in Air ($Le \neq 1$)	13
6. Stagnation Point Heat Transfer in Nitrogen ($Le = 1$)	14
7. Stagnation Heat-Transfer Parameter in Air or Nitrogen.	16
8. Summary of Various Data and Theories for Stagna- tion Point Heat Transfer in Partially Ionized Air after Rose and Stankevics (Ref. 10)	17

NOMENCLATURE

c_p	Specific heat at constant pressure, $\text{ft}^2/\text{sec}^2 - ^\circ\text{K}$
D	Binary diffusion coefficient
h	Enthalpy, ft^2/sec^2
h_D	Dissociation energy, ft^2/sec^2
h_E	Reference enthalpy, $2.119 \times 10^8 \text{ ft}^2/\text{sec}^2$
K	Constant (see Eq. (1))
k	Coefficient of thermal conductivity
Le	Lewis number, $\rho c_p D/k$
Nu/\sqrt{Re}	Heat-transfer parameter based on wall conditions (see Eq. (4))
p	Pressure, atm
\dot{q}	Heat-transfer rate, $\text{Btu}/\text{ft}^2\text{sec}$
R	Gas constant, $\text{ft}^2/\text{sec}^2 - ^\circ\text{K}$
r_n	Sphere (nose) radius, in.
S	Entropy, $\text{ft}^2/\text{sec}^2 - ^\circ\text{K}$
T	Temperature, $^\circ\text{K}$
U_∞	Free-stream velocity, $\sqrt{2h_0} \text{ ft}/\text{sec}$
u	Velocity at the edge of the boundary layer, ft/sec
x	Surface distance from the stagnation point
Z	Compressibility factor
μ	Dynamic viscosity, $\text{lb}_f\text{-sec}/\text{ft}^2$
μ_r	Reference viscosity (see Figs. 1 and 3)
ρ	Mass density in amagats where one amagat is the density of the perfect gas at one atmosphere pressure and 273.15°K

SUBSCRIPTS

a	Reference condition
o	Stagnation condition downstream of normal shock
w	Evaluated at the wall condition

All logarithms are to the base 10

REFERENCE QUANTITIES

$$T_a = 273.15^\circ\text{K}$$

$$p_a = 1 \text{ atm} = 2116.22 \text{ lb}_f/\text{ft}^2$$

	Air	Nitrogen
$\rho_a, \text{ lb}_f\text{-sec}^2/\text{ft}^4$	2.507542×10^{-3}	2.423609×10^{-3}
Z	1.00	1.00
$R, \text{ ft}^2/\text{sec}^2\text{ }^\circ\text{K}$	3089.67	3196.67

1.0 INTRODUCTION

Perhaps the most important and certainly one of the most investigated aerodynamic bodies is the sphere. For both drag and heat-transfer investigations, spheres and spherically nosed bodies have been extensively investigated throughout the history of aerodynamics. Since the sphere is so basic and amenable to theoretical and experimental investigations, it has been almost universally used as the nose of hypersonic, axisymmetric bodies. One can, in fact, use certain sphere data to predict results for other shapes. For these reasons heat-transfer rates at the stagnation point in equilibrium air and nitrogen are useful for wind tunnel and free-flight investigations.

Charts, developed by the von Kármán Gas Dynamics Facility (VKF), Arnold Engineering Development Center (AEDC), Air Force Systems Command (AFSC), are presented for the heat-transfer rate at the stagnation point of a sphere in air and nitrogen at chemical equilibrium. The calculations are based upon the modified formula of Fay and Riddell (Ref. 1) which is

$$\dot{q} \sqrt{r_n} = K (\rho_w \mu_w)^{0.41} (\rho_o \mu_o)^{0.4} (h_o - h_w) (p_o / \rho_o)^{0.25} \times \left\{ 1 + (Le - 1) \left[(h_{D,o} - h_{D,w}) / (h_o - h_w) \right] \right\}^{0.48} \quad (1)$$

where

$$K = 7.48873 \times 10^{-3} \text{ for nitrogen} \\ = 7.55274 \times 10^{-3} \text{ for air}$$

The term in braces differs from the one proposed by Fay and Riddell; they proposed $\{ 1 + (Le^{0.52} - 1) (h_D / h_o) \}$, whereas more recent work of Brokaw (Ref. 2) suggests the expression used above. * Equation (1) implies a Newtonian velocity gradient at the stagnation point. The adequacy of Newtonian theory has been shown for hypersonic free-stream conditions.

*For $T_w \sim 300^\circ K$ the two expressions in braces give almost identical results. However, for wall temperatures where dissociation may take place, Brokaw's formula should be more accurate.

Manuscript received June 1963.

2.0 THERMODYNAMIC AND TRANSPORT PROPERTIES

2.1 AIR

Thermodynamic properties were obtained from the data of Hilsenrath, Klein, and Woolley (Ref. 3) for air. These data were fit with thermodynamic surfaces by Lewis and Burgess (Ref. 4). For the range of the present calculations ($0.1 \leq p \leq 10$, $2000 \leq T \leq 15,000^\circ\text{K}$), the maximum errors were

	ρ	h/R	Z	T
Percent Error	2.42	1.96	0.75	2.24

The viscosity data of Hansen (Ref. 5) were used, and these data are shown on Fig. 1. The reference viscosity was

$$\mu_r = 3.0485 \times 10^{-6} \sqrt{T} \left[1 + (112/T) \right]^{-1}$$

The Lewis number was correlated by Cohen (Ref. 6) as a function of an enthalpy ratio and pressure. His correlation function, shown in Fig. 2, is independent of the pressure within 8 percent for $0.1 \leq p \leq 10$. The error is negligible at $p = 0.1$ and increases as p increases, becoming a maximum at $p = 10$.

If one assumes that high temperature air consists only of oxygen and nitrogen molecules and atoms, and that the dissociation of oxygen is complete before the nitrogen dissociation begins, then

$$h_D = \begin{cases} 1.830923 \times 10^6 (Z_o - 1), & 1.0 - Z \leq 1.21153 \\ 3.87295 \times 10^7 + 3.512878 \times 10^6 (Z_o - 1.21153), & Z > 1.21153 \end{cases}$$

For the present calculations, $T_w = 300^\circ\text{K}$ was assumed; thus, $h_{D,w} = 0$.

2.2 NITROGEN

The thermodynamic properties of equilibrium nitrogen were obtained from empirical surface fits of the data of Hilsenrath (Ref. 7) by Lewis and Burgess (Ref. 4). The maximum errors in the surface fits were

	ρ	h/R	Z	T
Percent Error	4.74*	3.62	0.86	5.09*

*The maximum errors for $T \leq 12,000^\circ\text{K}$ were 2.49 percent in ρ and 2.42 percent in T .

The viscosity data for nitrogen were taken from Ahtye and Peng (Ref. 8) and are shown plotted in Fig. 3.

3.0 RESULTS AND DISCUSSION

The stagnation heat-transfer rate for a sphere in equilibrium air with $Le = 1$ is shown in Fig. 4, where Z at $p = 0.1$ is also shown. The bends in the heat-transfer curve at the same pressure are easily correlated with changes in $Z(T; p = 0.1)$.

The Lewis number effect is shown in Fig. 5, where, for comparison, the $Le = 1$ results for three pressures are also shown. At the highest temperatures, the non-unity Lewis number effect on the heat-transfer rate is about 12 percent.

The stagnation heat-transfer in equilibrium nitrogen ($Le = 1$) is shown in Fig. 6.

A comparison of air and nitrogen with $Le = 1$ yields some interesting results. For definiteness, consider the pressure level $p_0 = 0.1$. The air and nitrogen results are almost identical up to $T_0 = 2500^\circ K$. Between 2500 and $6000^\circ K$, the heat-transfer rate in air is substantially higher (as much as 50 percent) than that in nitrogen. In this region the oxygen and nitrogen are dissociating. The oxygen dissociation is complete ($Z = 1.209$) at about $4200^\circ K$ at which temperature only about 1 percent of the nitrogen is dissociated.* At $6000^\circ K$ about 33 percent of the nitrogen is dissociated. This is indicated by an examination of both the air and nitrogen results. At about $6000^\circ K$ the heat-transfer rates are again almost equal. From 6000 to about $11,000^\circ K$, the heat-transfer rate in nitrogen is slightly higher (≤ 10 percent) than the air results. Between about 11,000 and $12,500^\circ K$ the results are once again almost identical. Above $12,500^\circ K$ there is another departure of results where the heat-transfer rate in nitrogen is again about 10 percent higher than in air. These comments generally apply at higher pressures; however, the temperatures quoted above will be changed as the pressure increases.

A comparison of the heat-transfer rates in air and nitrogen at $Le = 1$ as a function of stagnation enthalpy or free-stream velocity indicates almost identical results for all temperatures $2000 \leq T \leq 15,000^\circ K$ and pressures $p = 0.1$ to 10. This is in agreement with the assumption of Fay and Kemp (Ref. 9).

*This tends to substantiate the earlier assumption that the oxygen dissociation is complete before the nitrogen dissociation begins.

A correlation of the heat-transfer rate in terms of the parameter $\dot{q} \sqrt{r_n} / \sqrt{p_o}$ as a function of the free-stream velocity is shown in Fig. 7. The results in Figs. 4 and 6 for air and nitrogen are now reduced to a single curve. The solid line is accurate within ± 5 percent for both air and nitrogen in the pressure range $0.1 \leq p \leq 10$. The dashed curve represents an accuracy of ± 10 percent; however, it should be noted that in this latter region the Lewis number effects would presumably be greater than 10 percent (see Fig. 5).

The stagnation point heat-transfer rate in nitrogen based upon the Fay and Riddell theory was compared with the recent theory of Fay and Kemp and the shock-tube measurements of Rose and Stankevics (Ref. 10), and the results are shown in Fig. 8. In the present application of Fay and Riddell's theory in nitrogen for non-unity Lewis number, it was assumed that $Le = 0.6$ (constant) as was done by Fay and Kemp. Also, here it was assumed that

$$h_{D,0} = \begin{cases} 3.619269 \times 10^8 (Z_o - 1), & Z \leq 2 \\ 1.0743546 \times 10^8 (Z_o - 2) + 3.619269 \times 10^8, & Z > 2 \end{cases}$$

where the dissociation energy of nitrogen was 9.756 electron volts per molecule and the ionization energy was 14.48 electron volts per atom.

As previously noted by Rose and Stankevics (Ref. 10), the scatter in the experimental data was such that no theory could be judged superior. However, they indicated that their measurements clearly refuted the theory of Scala (Ref. 11) and the measurements reported by Warren (Ref. 12).

One interesting result is noted from the present calculation. If one simply multiplies the present Fay and Riddell results for Lewis number unity by 1.15, the results agree with Fay and Kemp's theory (equilibrium nitrogen, $Le = 0.6$) within about 5 percent for all velocities in $20,000 \leq U_\infty \leq 55,000$ ft/sec. It should be noted that the Fay and Kemp theory includes the effects of ionization which are neglected in the Fay and Riddell theory, and the ionization becomes significant at the higher velocities or temperatures (see, e. g., Fig. 4). Because of the present state of experimental data and the neglect here of radiation, an empirical modification of Fay and Riddell's theory may suffice for engineering estimates of the convective heat transfer. The modified formula becomes

$$\frac{Nu}{\sqrt{Re}} = \frac{(-\dot{q}_w) c_{p_w} x}{k_w (h_o - h_w)} \left(\frac{\rho_w u x}{\mu_w} \right)^{-1/2} = 0.77 \left(\frac{\rho_o \mu_o}{\rho_w \mu_w} \right)^{0.4} \quad (4)$$

4.0 CONCLUDING REMARKS

Fay and Riddell's theory has been used to compute the heat transfer to the stagnation point on a sphere. Air and nitrogen in chemical equilibrium were used in the calculations. Based upon these calculations the following results were obtained:

1. The effect of non-unity Lewis number was investigated in the air calculations. For the range of temperatures and pressures considered, the Lewis number effect was not greater than 15 percent.
2. There are rather large ranges of temperature and pressure over which the heat-transfer rates in air and nitrogen are essentially the same. However, oxygen dissociation strongly affects the heat-transfer rate in air between 2500 and 6000°K. However, the heat-transfer rates as a function of stagnation enthalpy and pressure in air and nitrogen ($Le = 1$) are almost identical (within 5 percent) in the range $2000 \leq T \leq 15,000^\circ K$.
3. Comparison of the Fay and Riddell theory (including only dissociation effects) with recent theories (such as that of Fay and Kemp which includes ionization effects) and shock-tube experimental measurements indicates that a simple empirical modification to the Fay and Riddell formula is adequate for engineering estimates.

REFERENCES

1. Fay, J. and Riddell, F. R. "Theory of Stagnation Point Heat Transfer in Dissociated Air." J. Aero. Sci., Vol. 25, No. 2, February 1958, pp. 73-85, 121.
2. Brokaw, R. S. "The Lewis Number." Progress in International Research on Thermodynamic and Transport Properties, ASME, 1962, pp. 271-278.
3. Hilsenrath, J., Klein, M., and Woolley, H. W. "Tables of Thermodynamic Properties of Air Including Dissociation and Ionization from 1500°K to 15,000°K." AEDC-TR-59-20, December 1959.

4. Lewis, C. H. and Burgess, E. G. "Empirical Equations for the Thermodynamic Properties of Air and Nitrogen to 15,000°K." AEDC-TDR-63-138, July 1963.
5. Hansen, C. H. "Approximation for the Thermodynamic and Transport Properties of High-Temperature Air." NACA TN 4150, March 1958.
6. Cohen, N. B. "Correlation Formulas and Tables of Density and Some Transport Properties of Equilibrium Dissociating Air for Use in Solutions of the Boundary-Layer Equations." NASA TN D-194, February 1960.
7. Hilsenrath, J. "A First Approximation to the Thermodynamic Properties of Nitrogen." National Bureau of Standards Preliminary Draft, May 1959 (Unpublished).
8. Ahtye, W. F. and Peng, T. C. "Approximations for the Thermodynamic and Transport Properties of High-Temperature Nitrogen with Shock-Tube Applications." NASA TN D-1303, July 1962.
9. Fay, J. R. and Kemp, N. "Theory of Stagnation Point Heat Transfer in a Partially Ionized Diatomic Gas." Paper No. 63-60 presented at the IAS 31st Annual Meeting, New York, January 1963.
10. Rose, P. H. and Stankevics, J. O. "Stagnation Point Heat Transfer Measurements in Partially Ionized Air." Paper No. 63-61, presented at the IAS 31st Annual Meeting, New York, January 1963.
11. Scala, S. M. and Warren, W. R. "Hypervelocity Stagnation Point Heat Transfer." ARS Journal, Vol. 32, 1962, p. 101.
12. Warren, W. R., Rogers, D. A., and Harris, C. J. "The Development of an Electrically Heated Shock Driven Test Facility." General Electric Company MSVD TIS Report R62SD37, April 1962.
13. Hoshizaki, H. "Heat Transfer in Planetary Atmospheres at Super-Satellite Speeds." ARS Journal, Vol. 32, 1962, p. 1544.
14. Rose, P. H. and Stark, W. I. "Stagnation Point Heat Transfer Measurements in Dissociated Air." Journal of the Aeronautical Sciences, Vol. 25, 1958, pp. 86-97.
15. Offenhartz, E., Weisblatt, H., and Flagg, R. F. "Stagnation Point Heat Transfer Measurements at Super-Satellite Speeds." J. Royal Aero Soc., Vol. 66, 1962, p. 63.

16. Pallone, A. and van Tassell, W. "The Effects of Ionization on Stagnation-Point Heat Transfer in Air and in Nitrogen." AVCO Tech Memo RAD TM 62-75, September 1962.
17. Cohen, N. B. "Boundary Layer Similar Solutions and Correlation Equations for Laminar Heat Transfer Distributions in Equilibrium Air at Velocities up to 41,000 Feet per Second." NASA Technical Report R-118, 1961.

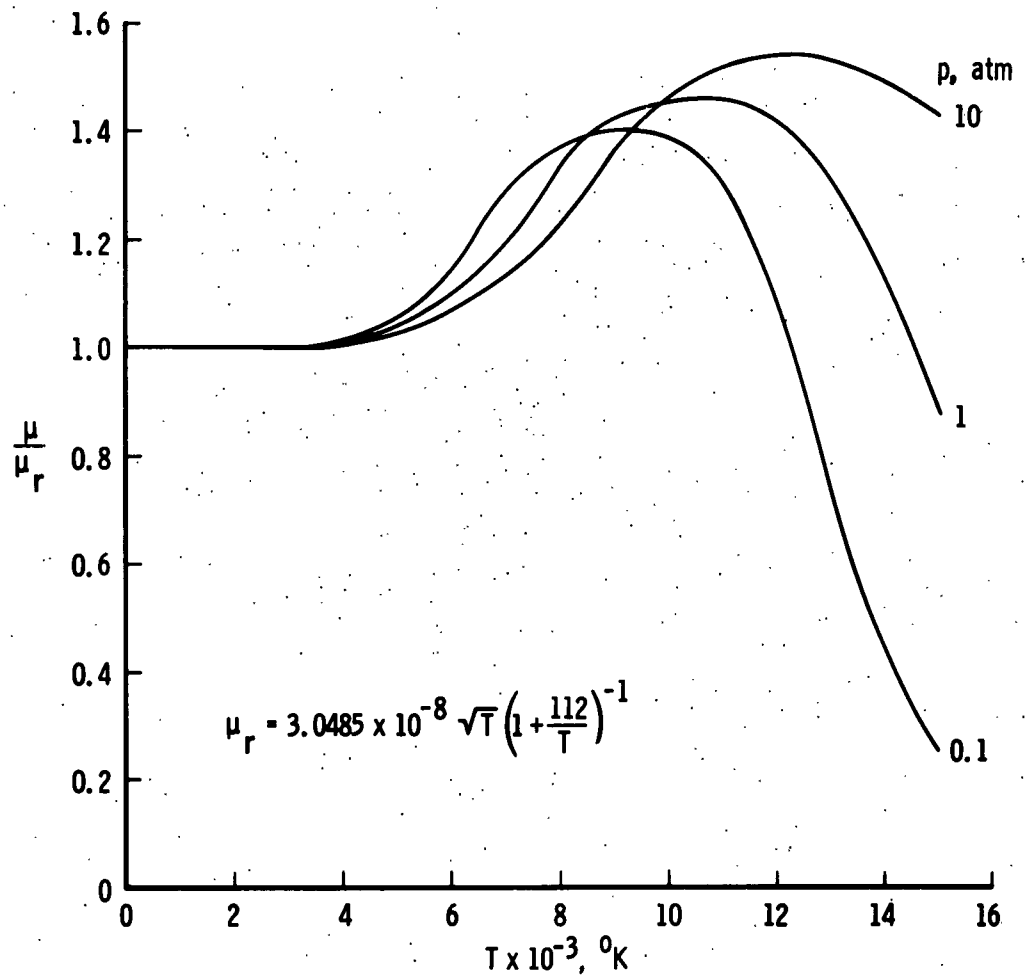


Fig. 1 Viscosity Ratio for Air

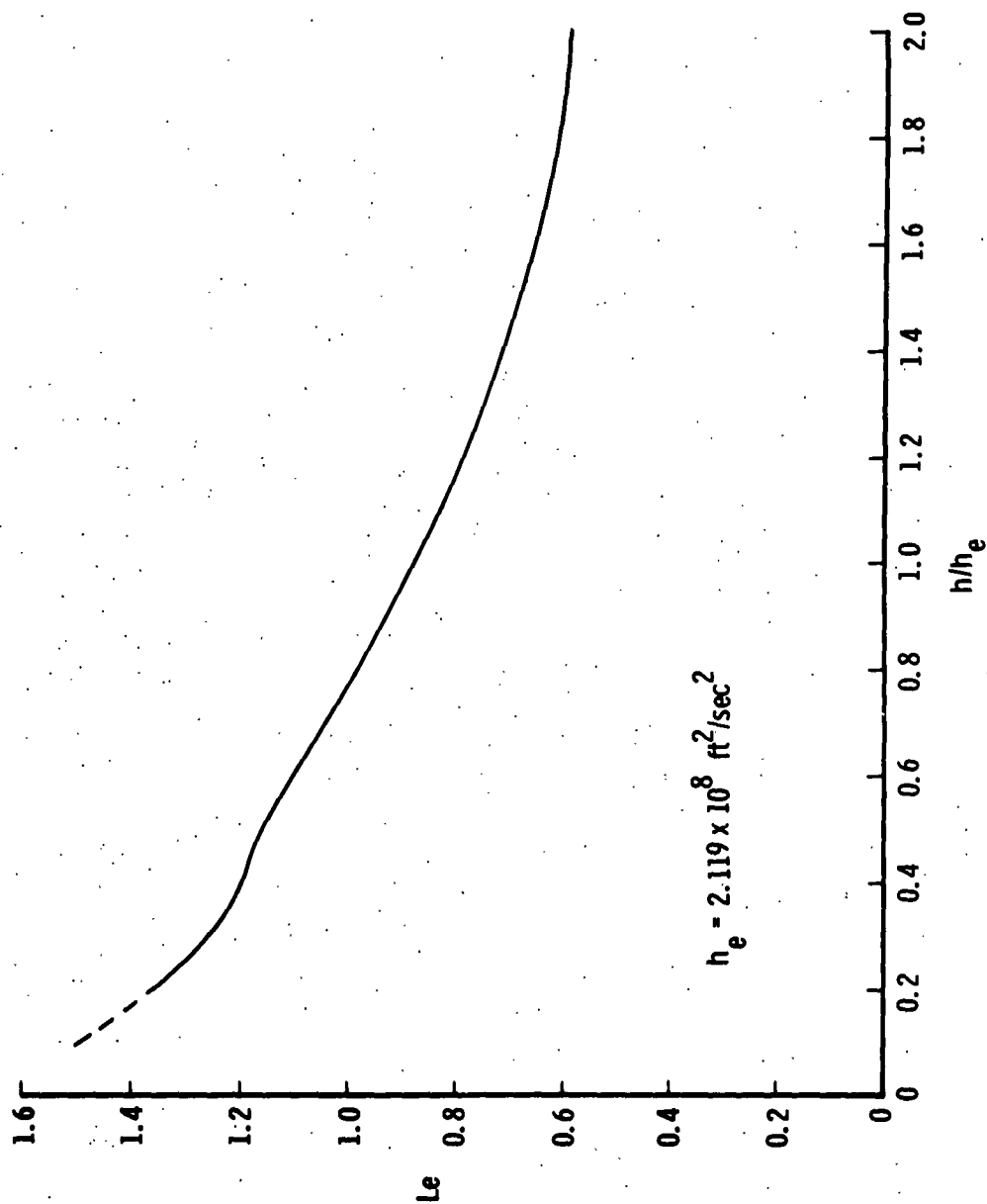


Fig. 2 Lewis Number Correlation for Air

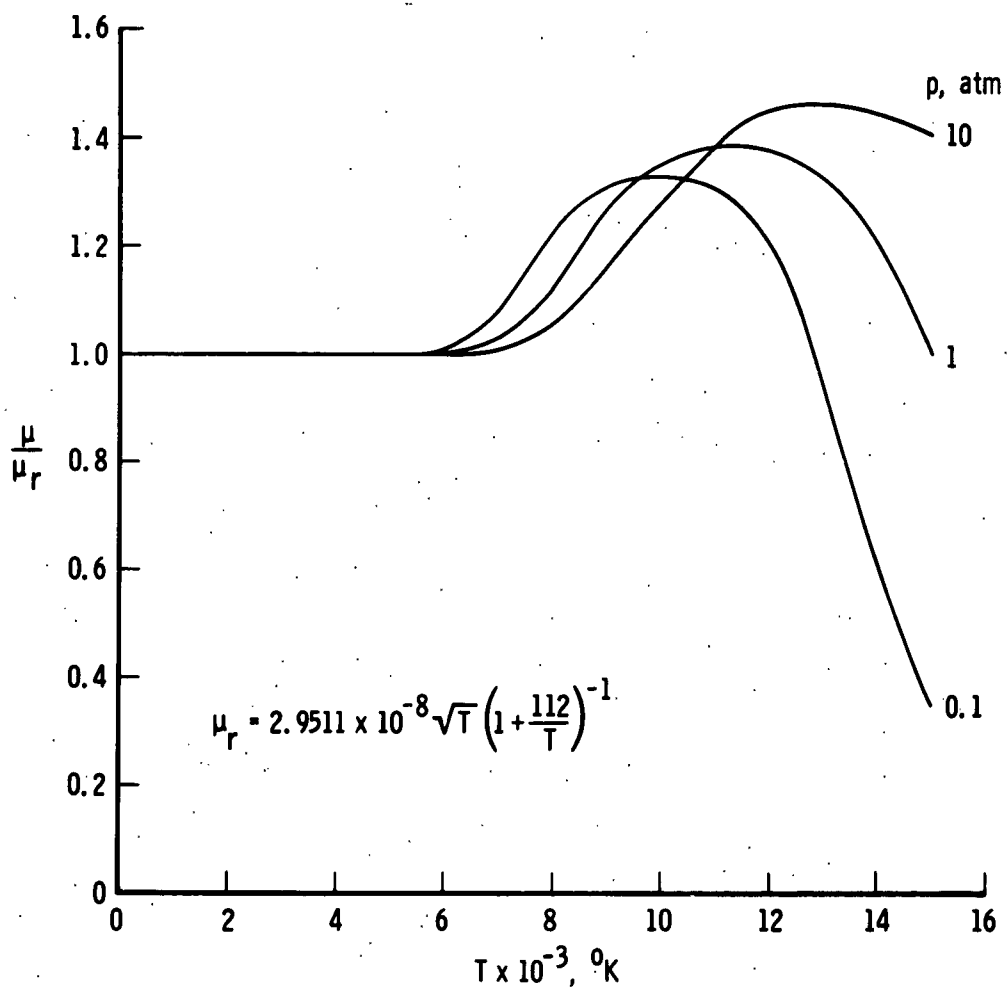
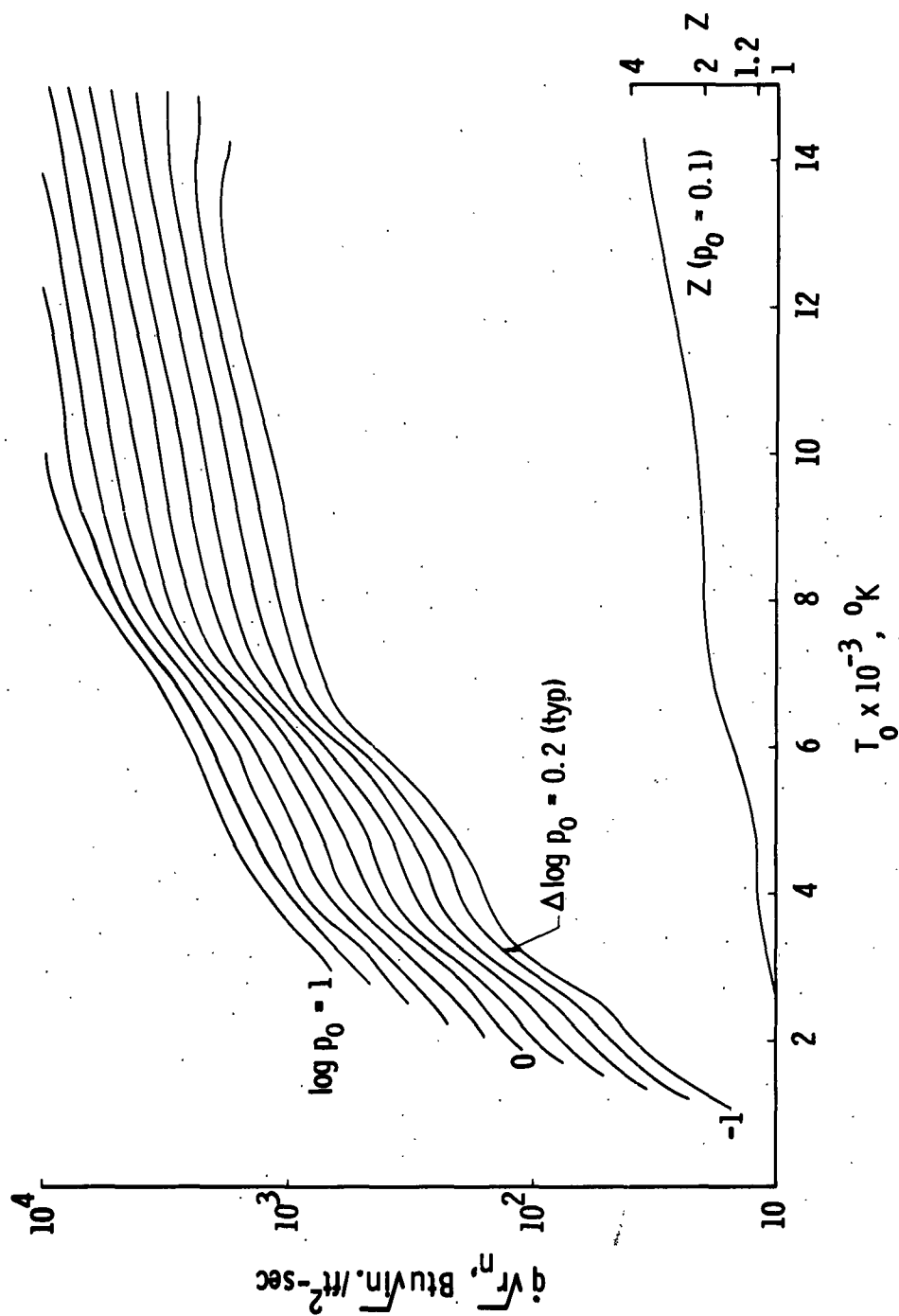


Fig. 3 Viscosity Ratio for Nitrogen

Fig. 4 Stagnation Point Heat Transfer in Air ($Le = 1$)

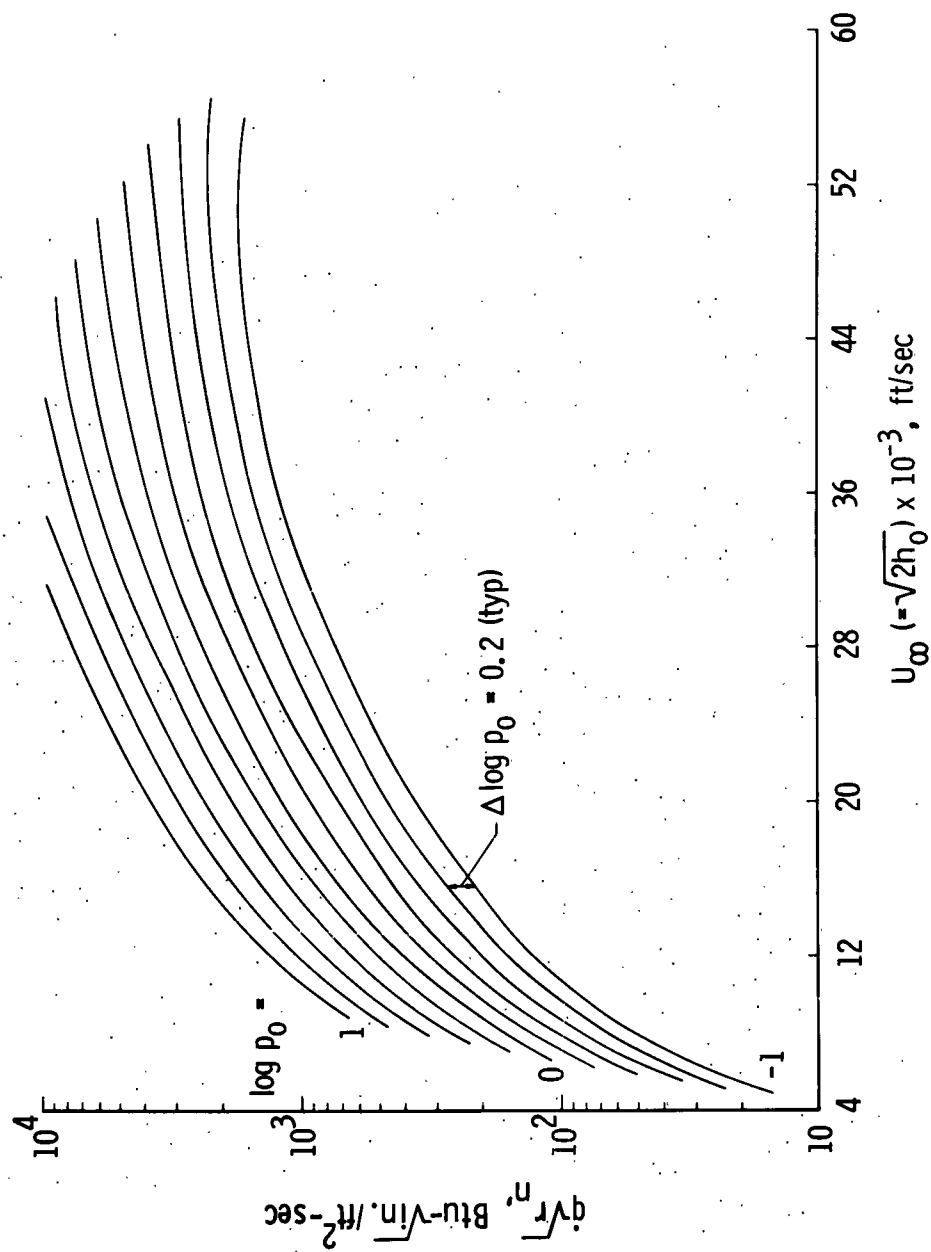
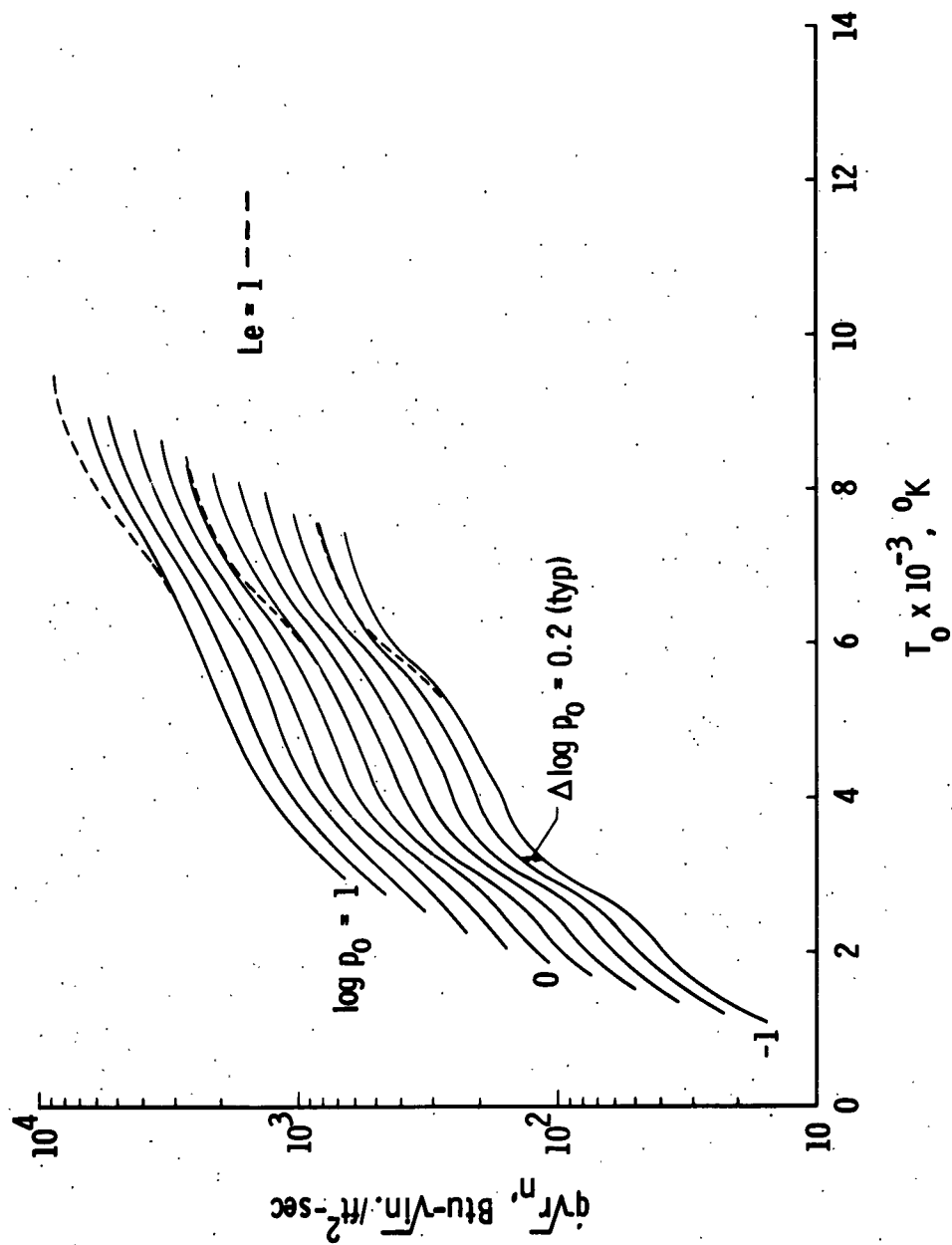
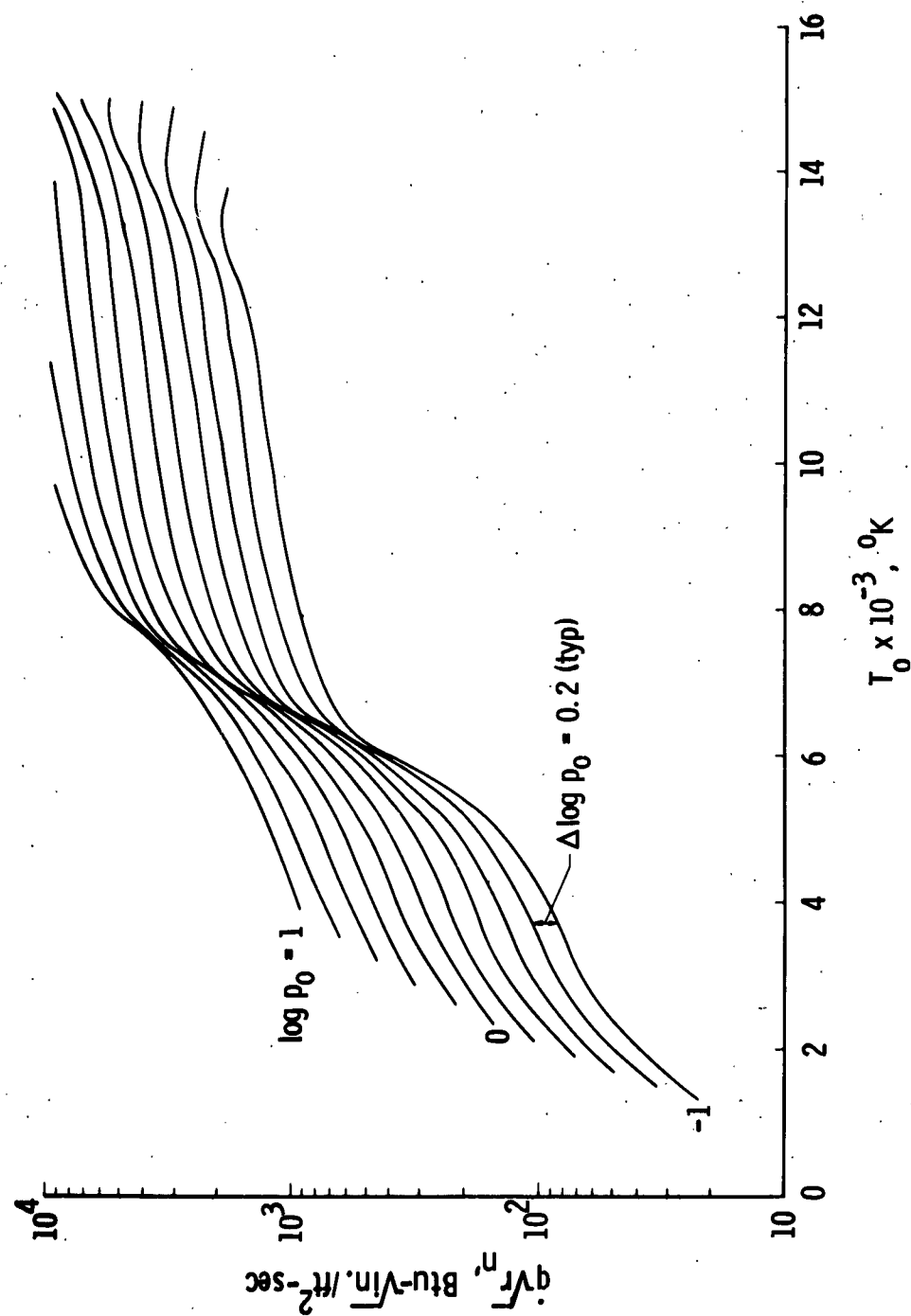


Fig. 4 Concluded

Fig. 5 Stagnation Point Heat Transfer in Air ($Le \neq 1$)

Fig. 6 Stagnation Point Heat Transfer in Nitrogen ($Le = 1$)

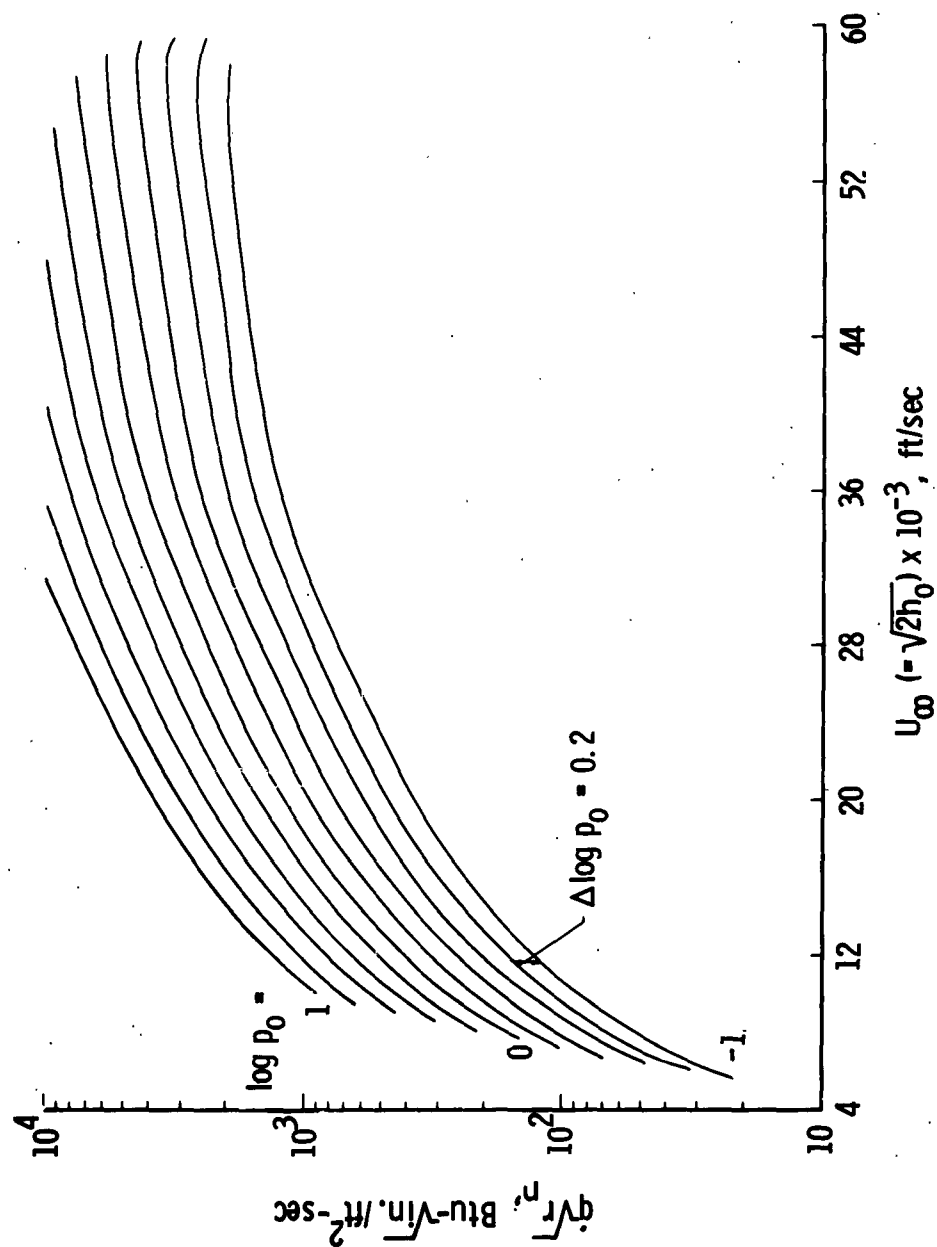


Fig. 6 Concluded

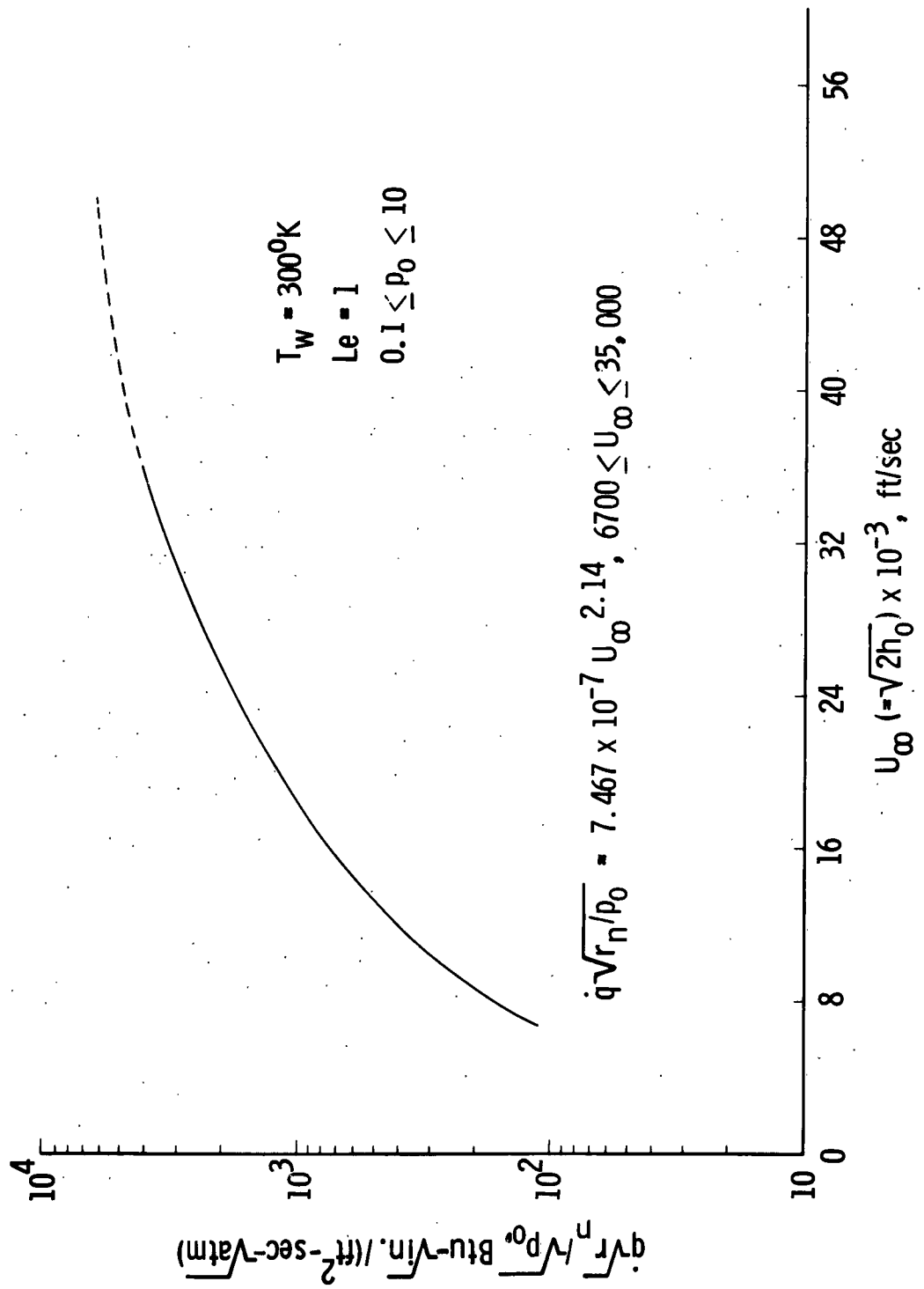


Fig. 7 Stagnation Heat-Transfer Parameter in Air or Nitrogen

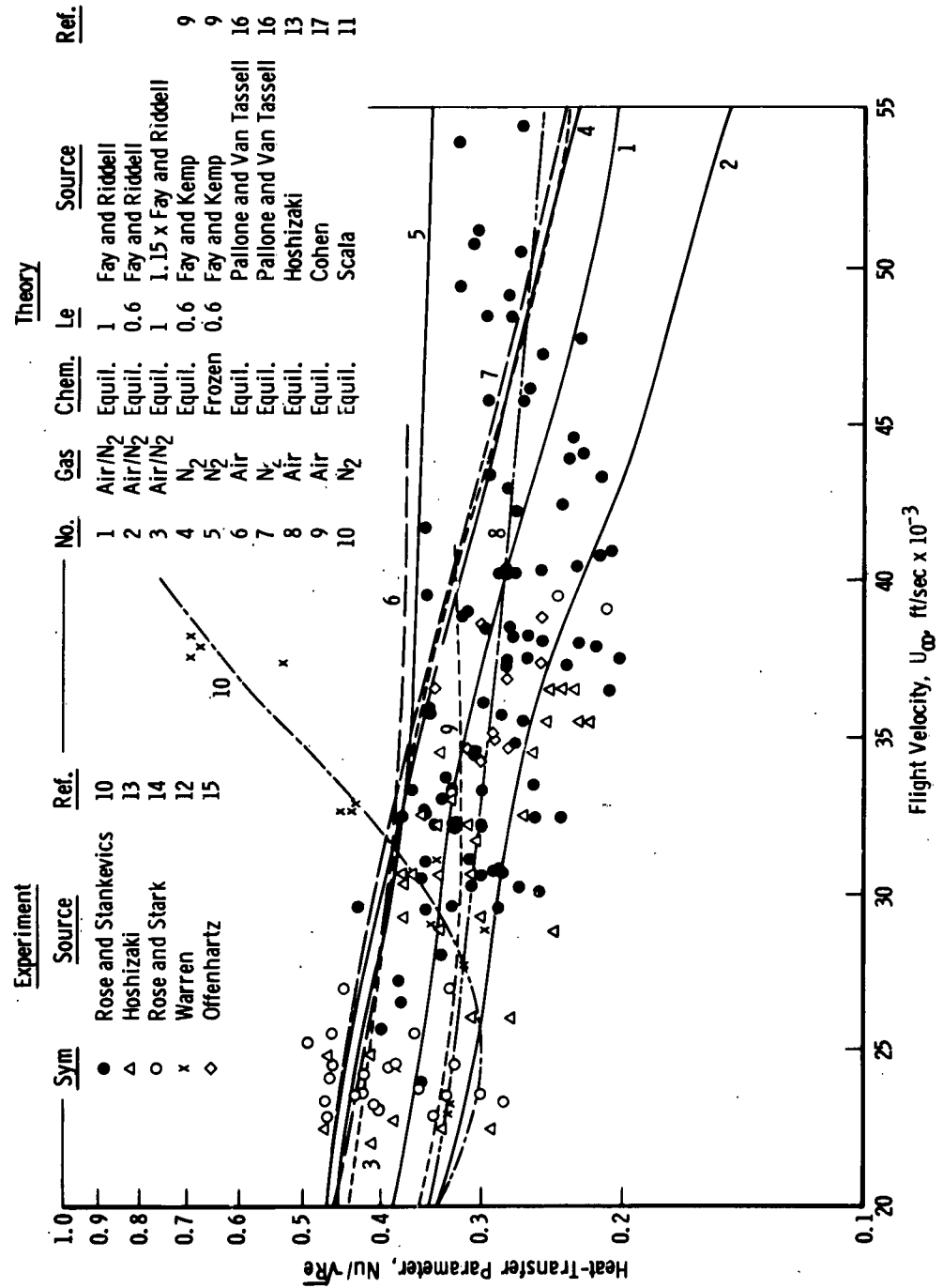


Fig. 8 Summary of Various Data and Theories for Stagnation Point Heat Transfer in Partially Ionized Air after Rose and Stankevics (Ref. 10).

<p>Arnold Engineering Development Center Arnold Air Force Station, Tennessee Rpt. No. AEDC-TDR-63-139. CHARTS OF SPHERE STAGNATION HEAT-TRANSFER RATE IN AIR AND NITROGEN AT HIGH TEMPERATURES. July 1963, 22 p. incl 17 refs., illus.</p> <p>Unclassified Report</p> <p>Fay and Riddell's theory for laminar heat transfer at the stagnation point is applied to the sphere in equilibrium air and nitrogen. Charts are presented of the heat-transfer rate as a function of stagnation temperature and pressure or free-stream velocity and stagnation pressure in the pressure range between 0.1 and 10 atm and in the temperature range from 2000 to 15,000°K. The effects of Lewis number are considered. Comparisons between Fay and Riddell's theory, the recent theory of Fay and Kemp, and shock-tube experiments are presented.</p>	<ol style="list-style-type: none"> 1. Air 2. Nitrogen 3. Heat transfer 4. Stagnation point 5. Spheres <ol style="list-style-type: none"> I. AFSC Program Area 040A II. Contract AF 40(600)-1000 III. ARO, Inc., Arnold AF Sta, Tenn. IV. C. H. Lewis and E. G. Burgess, III V. Available from OTS VI. In ASTIA Collection
<p>Arnold Engineering Development Center Arnold Air Force Station, Tennessee Rpt. No. AEDC-TDR-63-139. CHARTS OF SPHERE STAGNATION HEAT-TRANSFER RATE IN AIR AND NITROGEN AT HIGH TEMPERATURES. July 1963, 22 p. incl 17 refs., illus.</p> <p>Unclassified Report</p> <p>Fay and Riddell's theory for laminar heat transfer at the stagnation point is applied to the sphere in equilibrium air and nitrogen. Charts are presented of the heat-transfer rate as a function of stagnation temperature and pressure or free-stream velocity and stagnation pressure in the pressure range between 0.1 and 10 atm and in the temperature range from 2000 to 15,000°K. The effects of Lewis number are considered. Comparisons between Fay and Riddell's theory, the recent theory of Fay and Kemp, and shock-tube experiments are presented.</p>	<ol style="list-style-type: none"> 1. Air 2. Nitrogen 3. Heat transfer 4. Stagnation point 5. Spheres <ol style="list-style-type: none"> I. AFSC Program Area 040A II. Contract AF 40(600)-1000 III. ARO, Inc., Arnold AF Sta, Tenn. IV. C. H. Lewis and E. G. Burgess, III V. Available from OTS VI. In ASTIA Collection